قائمة الاسئلة

الكيمياء التحليلية الصيدلية 1-مستوى ثاني علوم صيدلانية-درجة الامتحان 75 درجة

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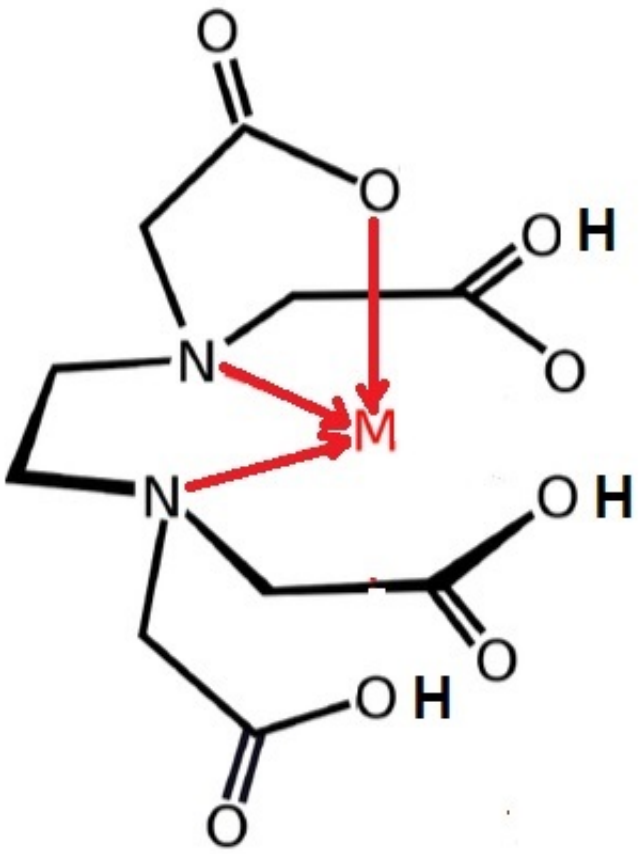
1) Complexometric formation titration involves:
   1) - (a) Acid base substances.
   2) - (b) Precipitating agent.
   3) + (c) Complexing agent.
   4) - (d) All the mentioned.
2) In complexometry:
   1) - (a) Double bonds are formed only.
   2) - (b) Hydrogen bonds only.
   3) + (c) Coordinated bonds are formed.
   4) - (d) All the mentioned.
3) Administration of tetracycline and milk leading to:
   1) - (a) Acid base reaction.
   2) - (b) Precipitation reaction.
   3) + (c) Complexation reaction.
   4) - (d) All the mentioned.
4) The complex in the figure , according to the complexing agent can be called

Q4.



   1) - (a) Monodentate.

2) + (b)Tri-dentate complex.
3) - (c) Tri-nuclear complex.
4) - (d) All the mentioned.

5) One of the following types of indicators used in detection of the end point in complexometry
1) - (a) Coloring indicators.
2) - (b) Adsorptive indicators.
3) + (c) Metallochromic indicators.
4) - (d) All the mentioned.

6) When EDTA used in complexometry:
1) - (a) The end point depends on pH.
2) - (b) The end point depends on EDTA volume.
3) - (c) The titration curve between EDTA volume and the p Metal.
4) + (d) All the mentioned.

7) EDTA is
1) - (a) Selective complexing agent.
2) + (b) Non-selective complexing agent.
3) - (c) (a&b).
4) - (d) None of the mentioned.

8) During analysis of water hardness due to Ca+2 and Mg +2, total hardness at pH
1) + (a) 10.
2) - (b) 12.
3) - (c) Neutral.
4) - (d) All the mentioned.

9) During analysis of water hardness due to (Ca+2 and Mg +2), Ca +2 hardness is determined at pH
1) - (a) 10.
2) + (b) 12.
3) - (c) Neutral.
4) - (d) All the mentioned.

10) To improve selectivity of EDTA during complexometry ,we can
1) - (a) Adjust the pH.
2) - (b) Add masking agent.
3) + (c) (a&b).
4) - (d) None of the mentioned.

11) Calcium gluconate and magnesium salt can be analyzed and determined using
1) - (a) Acid- base titration.
2) - (b) Precipitation titration.
3) + (c) Complex formation titration.
4) - (d) All the mentioned.

12) One of the following types of indicators used in detection of the end point in precipitation titration
1) - (a) Coloring indicators.
2) - (b) Metallochromic indicators.
3) - (c) Complex formation indicators..
4) + (d) (a&c).

13) Precipitation of a precipitate is formed when the concentration product of the reactants of the precipitate is
1) + (a) ≥ Ksp ppt value.
2) - (b) ≤ Ksp ppt value.
3) - (c) > Ksp ppt value.
4) - (d) All the mentioned.

14) In precipitation titration , we use

1) - (a) Acid - base.
2) - (b) Complexing agents.
3) + (c) Precipitating agent.
4) - (d) All the mentioned.

15) When we use Ag as precipitating agent , this method is called
1) - (a) Complexometry.
2) + (b) Argentometry.
3) - (c) Iodometry.
4) - (d) All the mentioned.

16) The titration curve of precipitation titration is a plot between
1) - (a) pH and the titrant volume.
2) + (b) p sample and titrant volume.
3) - (c) p metal and EDTA volume.
4) - (d) All the mentioned.

17) In Mohr's method, one of the following indicators is used:
1) - (a) Adsorptive indicators as Eosin.
2) - (b) Complex formation indicator as ferric ammonium sulphate.
3) + (c) Coloring indicator as KMnO4.
4) - (d) All the mentioned.

18) In Volhard's method , one of the following indicators is used:
1) - (a) Adsorptive indicators as Eosin.
2) + (b) Complex formation indicator as ferric ammonium sulphate.
3) - (c) Coloring indicator as KMnO4.
4) - (d) All the mentioned.

19) In Fajan's method, one of the following indicators is used:
1) + (a) Adsorptive indicators as Eosin.
2) - (b) Complex formation indicator as ferric ammonium sulphate.
3) - (c) Coloring indicator as KMnO4.
4) - (d) All the mentioned.

20) One of the following methods is a non-direct (back or indirect) titration method:
1) - (a) Mohr's method.
2) + (b) Volhard's method.
3) - (c) Fajan's method.
4) - (d) All the mentioned.

21) When you have (Cl- , Br- and I-) sample mixture , you will have:
1) - (a) One end point.
2) - (b) Two end points.
3) + (c) Three end points.
4) - (d) None of the mentioned.

22) If you know that Ksp AgCl =1x10 -6 M, Ksp AgBr = 1x 10 -3 M and Ksp AgI =1x 10 -8 M, the precipitation order as follow
1) + (a) I- then Cl- and finally Br-.
2) - (b) Cl- then I- and finally Br-.
3) - (c) Br- then Cl- and finally I-.
4) - (d) All the mentioned.

23) If you know that Ksp AgCl = 1x 10 6 M , [Cl-] = 1x 10 2 M and [Ag+] = 1x10 3 M ,so
1) - (a) AgCl is precipitated .
2) + (b) AgCl is not precipitated.
3) - (c) (a&b).

4) - (d) None of the mentioned.

24) Adsorption process meaning:
1) + (a) Adhesion of a substance molecules on the surfaces of another substance.
2) - (b) A penetration of a substance molecules into another substance molecules.
3) - (c) (a&b).
4) - (d) None of the mentioned.

25) The most accurate and recent precipitation method is:
1) - (a) Mohr's method.
2) - (b) Volhard's method.
3) + (c) Fajan's method.
4) - (d) All the mentioned.

26) Thiamine chloride can be determined using:
1) - (a) Acid base titration.
2) + (b) Precipitation titration.
3) - (c) Complex formation titration.
4) - (d) All the mentioned.

27) Silver nitrate (external antiseptic) can be analysed using
1) - (a) Acid base titration.
2) + (b) Precipitation titration.
3) - (c) Complex formation titration.
4) - (d) All the mentioned.

28) From applcations of precipitation titration:
1) - (a) Determination of HCl.
2) - (b) Determination of Calcium.
3) + (c) Determination of Normal saline solutions.
4) - (d) All the mentioned.

29) The indicators always play important role in detection of the end point :
1) - (a) Before end point.
2) + (b) At the end point.
3) - (c) After end point.
4) - (d) None of the mentioned.

30) An example of metallochromic indicator:
1) - (a) Methyl orange.
2) - (b) Eosine.
3) - (c) Phenolphthalein.
4) + (d) Eriochrome black T.

31) In pharmaceutical analysis we carry out standard test in order to:
1) - (a) Reduce errors.
2) - (b) To calculate sample concentration.
3) - (c) To compare between standard and sample result.
4) + (d) (b&c).

32) Residual back titration can be done when , the sample acidity and alkalinity is weak:
1) + (a) True.
2) - (b) False.

33) In Residual back titration :
1) - (a) One standard solution used only.
2) + (b) Two standard solutions can be used.
3) - (c) Three standard solutions can be used.
4) - (d) None of the mentioned.

34) In Residual back titration of NH4Cl , we use:
1) - (a) The first standard solution is HCl and in burette we use the second standard solution of NaOH.
2) + (b) The first standard solution is NaOH and in burette we use the second standard solution of HCl.
3) - (c) (a&b).
4) - (d) None of the mentioned.

35) We can use non-aqueous acid base titration , when:
1) - (a) The sample soluble in water.
2) - (b) The sample insoluble in water.
3) - (c) The ionization constant of the sample in water 1x10-8.
4) + (d) (b&c).

36) One of the following can be determined using non-aqueous acid base titration:
1) - (a) HCl.
2) - (b) NaOH.
3) + (c) Phenobarbitone.
4) - (d) None of the mentioned.

37) In determination of aniline ,we can use:
1) - (a) Precipitation titration.
2) - (b) Complexometry.
3) - (c) Argentometry.
4) + (d) Non-aqueous acid base titration.

38) Protophilic solvents are:
1) - (a) Acidic in nature.
2) - (b) Basic in nature.
3) - (c) Have tendency to accept proton.
4) + (d) (b&c).

39) Aprotic solvents are:
1) - (a) Have not H+.
2) - (b) As CCl4.
3) - (c) Non-aqueous solvents.
4) + (d) All the mentioned.

40) The solvents in non-aqueous titrations used as media only and not enter in the reaction:
1) + (a) True.
2) - (b) False.

41) Acid base titration curve is a plot between:
1) + (a) pH and titrant volume.
2) - (b) p sample and titrant volume.
3) - (c) p metal and EDTA volume.
4) - (d) All the mentioned.

42) One of the following indicators used in acid-base titration:
1) - (a) Adsorptive indicators as Eosin.
2) - (b) Complex formation indicator as ferric ammonium sulphate.
3) + (c) Coloring indicator as methyl orange.
4) - (d) All the mentioned.

43) Acetyl salicylic acid can be determined by:
1) - (a) Precipitation titration.
2) - (b) Complexometry.
3) + (c) Acid base titration.
4) - (d) All the mentioned.

44) When you titrate H2SO4 with NaOH , you have

1) - (a) One end point.
2) + (b) Two end points.
3) - (c) Three end points.
4) - (d) None of the mentioned.

45) When acetyl salicylic acid is used with vitamin C (Ascorbic acid) ,this leading to:
1) + (a) Increased drug absorption and decreased drug excretion.
2) - (b) Decreased drug absorption and increased drug excretion.
3) - (c) The ionized form is more than the unionized form of the drug.
4) - (d) All the mentioned.

46) Most of basic drugs absorbed highly from:
1) - (a) The stomach.
2) + (b) The intestine.
3) - (c) Blood brain barrier.
4) - (d) All the mentioned.

47) Most of acidic drugs absorbed highly from:
1) + (a) The stomach.
2) - (b) The intestine.
3) - (c) Blood brain barrier.
4) - (d) All the mentioned.

48) pH of a sample meaning -log[H+]:
1) + (a) True.
2) - (b) False.

49) When the [H+] = 1x 10 -6 M, this meaning that :
1) + (a) pH =6.
2) - (b) pH =-6.
3) - (c) pH=0.
4) - (d) None of the mentioned.

50) Controlling pH value constant by:
1) + (a) Buffer solution.
2) - (b) Serial dillution.
3) - (c) Concentrated solution.
4) - (d) None of the mentioned.